

Name: _____

Just How Small is an Atom

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1. What is in between the nucleus and the electrons in an atom?

- a. A jello-like goo
- b. Protons and Neutrons
- c. A mushy substance similar to pudding
- d. Empty space

2. Describe the size and scale of an atom. Explain the relative size of the nucleus in relation to the size of the entire atom.

3. What is contained in the nucleus of an atom?

- a. Protons and neutrons
- b. Electrons
- c. Neutrons only
- d. The DNA and chromosomes

4. Come up with your own analogy as to the size of an atom. Do some proportions to determine how you could describe how small an atom is.

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5. The density of a typical nucleus is equal to....

- a. Similar to 6.2 billion cars in a one cubic foot box
- b. Similar to the density of lead
- c. Similar to the density of water (1 g/mL)
- d. Similar to the density of an aircraft carrier.

6. The number of atoms in a grapefruit would be equivalent to

- a. The number of blueberries in the Sun
- b. The number of blueberries in a house
- c. The number of blueberries that would fit into the world
- d. The number of blueberries in the Ocean

7. If you had an atom that was blown up to the size of a blueberry, the nucleus would be the size of....

- a. It would be invisible
- b. It would be the size of a grain of sand
- c. It would be the size of a marble
- d. It would be the size of golf ball